

- (1) Water has dual character. How?
- (2) What is the limitation of Arrhenius concept of Acid and base?

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Some Important points about Conjugate acid and base

- Rule (1) Stronger Acid has weaker CB and vice-versa
- Rule (2) Stronger base has weaker conjugate Acid and vice-versa.

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Lewis concept of Acid and bases →

Arrhenius and Bronsted-Lowry Concepts of

Acid and bases are limited to hydrogen and oxygen containing substances are acid and bases

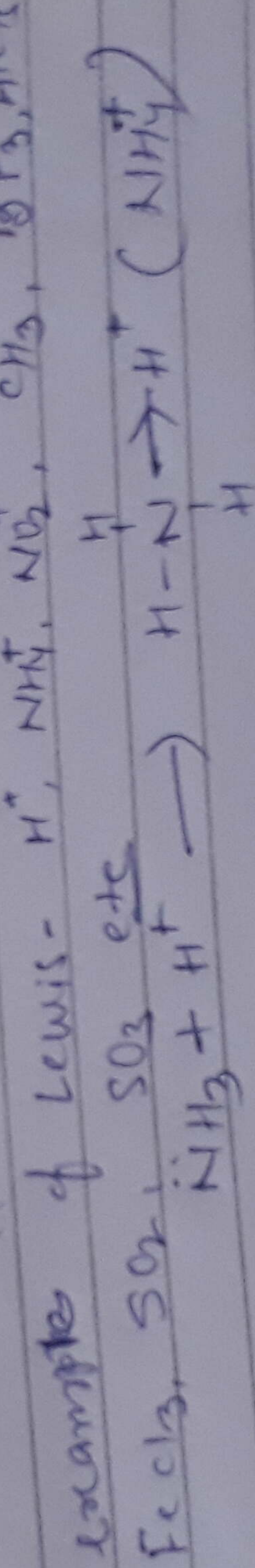
But G.N. Lewis in 1923 generalised the term of acid and bases which is based on electron.

therefore Lewis concept is also called electronic concept of acids and bases.

According to this concept —

Acid is a substance (an atom, molecule, or ion) which can accept a pair of electrons and base is a substance (an atom, molecule, or ion) which can donate a pair of electrons.

i.e. Electron pair acceptor — Acid
 " " donor — Base.



The Neutralization ~~is~~ ~~between~~ Acid and Base means the formation of co-ordinate bond

